



*Supplement of*

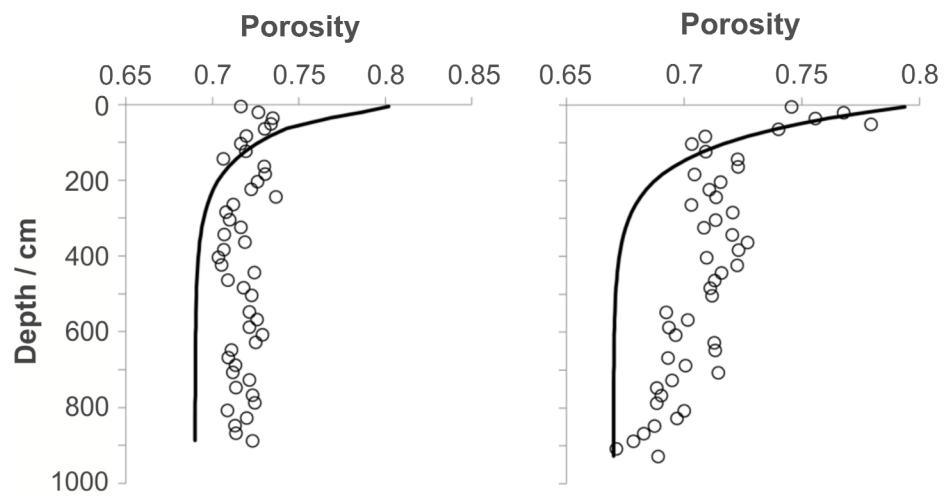
**Impact of small-scale disturbances on geochemical conditions,  
biogeochemical processes and element fluxes in surface sediments  
of the eastern Clarion–Clipperton Zone, Pacific Ocean**

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**Figure S1:** Measured porosity (dots) and numerical approximation (solid line) for the BGR-RA (left) and IOM (right) sites. See Table S2 for fitting details.



**Table S1:** Electron-equivalent redox reactions used in the transient transport-reaction model for the 1-day-old EBS disturbance in the BGR-RA area and for the 20-year-old IOM-BIE disturbance in the IOM area.

Reaction name	Reaction	Redox reaction
Aerobic respiration	R <sub>1</sub>	$(CH_2O)(NH_4^+)_\frac{16}{106} + O_2 \rightarrow CO_2 + \frac{16}{106} NH_4^+ + H_2O$
Heterotrophic denitrification	R <sub>2</sub>	$5(CH_2O)(NH_4^+)_\frac{16}{106} + 4NO_3^- + 4H^+ \rightarrow 2N_2 + 5CO_2 + 5\frac{16}{106} NH_4^+ + 7H_2O$
Dissimilatory Mn(IV) reduction	R <sub>3</sub>	$(CH_2O)(NH_4^+)_\frac{16}{106} + 2MnO_2 + 4H^+ \rightarrow 2Mn^{2+} + CO_2 + \frac{16}{106} NH_4^+ + 3H_2O$
Mn <sup>2+</sup> oxidation	R <sub>4</sub>	$2Mn^{2+} + O_2 + 2H_2O \rightarrow 2MnO_2 + 4H^+$
Nitrification	R <sub>5</sub>	$NH_4^+ + 2O_2 \rightarrow NO_3^- + 2H^+ + H_2O$
Mn-anammox	R <sub>6</sub>	$3MnO_2 + 2NH_4^+ + 4H^+ \rightarrow 3Mn^{2+} + N_2 + 6H_2O$

**Table S2:** Species, boundary conditions and fitted parameter values used in the transient transport-reaction models for the 1-day-old EBS disturbance in the BGR-RA area and the 20-year-old IOM-BIE disturbance in the IOM area.

	Symbol	Unit	BGR	IOM
<b><i>Species and boundary conditions</i></b>				
Porosity at SWI	$\varphi_0$		0.82	0.8
Porosity at compacted depth	$\varphi_\infty$		0.69	0.67
Sedimentation rate	$\omega_i$	cm kyr <sup>-1</sup>	0.65	1.15
Oxygen <sub>bw</sub>	$O_2$	µM	120	150
Ammonium <sub>bw</sub>	$NH_4^+$	µM	1	1
Nitrate <sub>bw</sub>	$NO_3^-$	µM	50	38
Dissolved reduced manganese <sub>bw</sub>	$Mn^{2+}$	µM	1	1
<b><i>Fitted parameters</i></b>				
Sediment thickness		m	10	10
Removed sediment thickness		cm	10	7
Labile C <sub>org</sub>	$TOC_1$	mol m <sup>-2</sup> yr <sup>-1</sup>	6.0E-02	4.5E-02
Metabolizable C <sub>org</sub>	$TOC_2$	mol m <sup>-2</sup> yr <sup>-1</sup>	1.2E-04	1.1E-03
Refractory C <sub>org</sub>	$TOC_3$	mol m <sup>-2</sup> yr <sup>-1</sup>	5.0E-04	7.8E-04
Oxygen <sub>bas</sub>	$O_2$	µM	18	55
Ammonium <sub>bas</sub>	$NH_4^+$	µM	1	1
Nitrate <sub>bas</sub>	$NO_3^-$	µM	19	30
Dissolved reduced manganese <sub>bas</sub>	$Mn^{2+}$	µM	1	1
1 <sup>st</sup> order deg. coeff. TOC <sub>1</sub>	$\sigma_1$	yr <sup>-1</sup>	1.0E-03	1.0E-02
1 <sup>st</sup> order deg. coeff. TOC <sub>2</sub>	$\sigma_2$	yr <sup>-1</sup>	1.0E-06	5.5E-06
1 <sup>st</sup> order deg. coeff. TOC <sub>3</sub>	$\sigma_3$	yr <sup>-1</sup>	2.0E-09	2.5E-09
Bioturbation coefficient	$B_0$	cm <sup>2</sup> yr <sup>-1</sup>	0.5	0.2
Biomixing half depth	$z_{mix}$	cm	7.0	7.0
Biomixing attenuation	$zz_{att}$	cm	0.1	0.1
Bioirrigation coefficient	$\alpha_0$	yr <sup>-1</sup>	0.65	2.0
O <sub>2</sub> inhibition concentration for R <sub>1</sub>	$h_1$	µM	0.008	0.006
NO <sub>3</sub> <sup>-</sup> inhibition concentration for R <sub>2</sub>	$h_2$	µM	45	25
R <sub>4</sub> rate constant	$k_4$	µM <sup>-1</sup> yr <sup>-1</sup>	0.1	0.1
R <sub>5</sub> rate constant	$k_5$	µM <sup>-1</sup> yr <sup>-1</sup>	0.005	0.1
R <sub>6</sub> rate constant	$k_6$	µM <sup>-1</sup> yr <sup>-1</sup>	0.001	0.001

**Table S3:** Electron-equivalent redox reactions and associated expressions used in the numerical diagenetic model.

Reaction name	Reaction	Rate expression
Aerobic respiration	R <sub>1</sub>	$(\sigma_1 C_{TOC1} + \sigma_2 C_{TOC2} + \sigma_3 C_{TOC3}) \frac{C_{O_2}}{C_{O_2} + h_1}$
Heterotrophic denitrification	R <sub>2</sub>	$(\sigma_1 C_{TOC1} + \sigma_2 C_{TOC2} + \sigma_3 C_{TOC3}) \gamma \frac{C_{NO_3^-}}{C_{NO_3^-} + h_2}$
Dissimilatory Mn(IV) reduction	R <sub>3</sub>	$(\sigma_1 C_{TOC1} + \sigma_2 C_{TOC2} + \sigma_3 C_{TOC3}) \gamma \frac{h_2}{C_{NO_3^-} + h_2}$
Mn <sup>2+</sup> oxidation	R <sub>4</sub>	$k_4 C_{O_2} C_{Mn^{2+}}$
Nitrification	R <sub>5</sub>	$k_5 C_{O_2} C_{NH_4^+}$
Mn-annamox	R <sub>6</sub>	$k_6 C_{NH_4^+} C_{MnO_2} \gamma \frac{h_2}{C_{NO_3^-} + h_2}$

$$\gamma = \frac{h_1}{(h_1 + C_{O_2})}$$